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- (b) Chloroform (CH_3Cl) molecules are polar (why?). Dispersion and dipole-dipole forces will be present.
- (c) Phosphorus trifluoride (PF_3) molecules are polar. Dispersion and dipole-dipole forces will be present.
- (d) Sodium chloride ($NaCl$) is an ionic compound. Ion-ion (and dispersion) forces will be present.
- (e) Carbon disulfide (CS_2) molecules are nonpolar. Only dispersion forces will be present.

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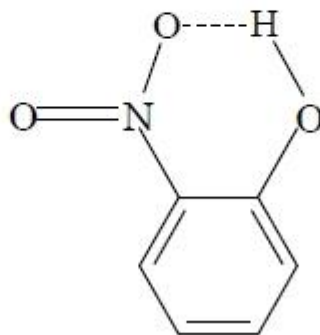
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